

The 'Lime Cycle'



Cut out and organise to form the 'Lime Cycle'

Calcium hydroxide reverts
back to Calcium carbonate,
releasing water
 $\text{Ca(OH)}_2 + \text{CO}_2 = \text{CaCO}_3 + \text{H}_2\text{O}$

QUICKLIME
Calcium Oxide
($\text{CaO} + \text{CO}_2$)

LIMESTONE
Calcium carbonate
(CaCO_3)

Slaking
(water added)
 $\text{CaO} + \text{H}_2\text{O}$

SLAKED LIME
Calcium hydroxide
 $\text{CaO} + \text{H}_2\text{O} = \text{Ca(OH)}_2$

hardening
Slaked lime
absorbs CO_2
from the air

calcined
(roasted)
(heat added)

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